

## *Chapter 10 Chemical Quantities Test Answers*







### Chapter 10 Chemical Quantities Test

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CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The mole was founded by a scientist named Avogadro, and he decided to use the

### CHAPTER 10: Chemical Quantities

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### Chapter 10 Chemical Quantities Chapter Test A Answers

Chapter 10 (Chemical Quantities) Test Study Guide The mole is the SI unit used to measure the number of representative particles in a substance. A representative particle can be an atom, an ion, or a molecule, depending upon the way a substance commonly exists. Avogadro's number ( $6.02 \times 10^{23}$ ) is the number of representative particles of a substance present in 1 mole of that substance.

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Chapter 10 - Chemical Quantities. Jennie L. Borders. Section 10.1 - The Mole: A Measurement of Matter. You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is  $6.02 \times 10^{23}$  representative particles of that substance.

### Chapter 10 - Chemical Quantities

Chapter 10 "Chemical Quantities" Vocab. Tools. Copy this to my account; E-mail to a friend ... Use these activities to learn the vocabulary presented in this chapter. A B; percent composition: a description of the relative amounts of each element in a compound: empirical formula: the lowest whole- number ratio of the atoms of the elements in a ...

### Quia - Chapter 10 "Chemical Quantities" Vocab

Chapter 10 Chemical Quantities 243 Section Review Objectives • Convert the mass of a substance to the number of moles of a substance, and the number of moles of a substance to mass • Calculate the volume of a quantity of gas at STP Vocabulary • Avogadro's hypothesis • standard temperature and pressure (STP) • molar volume Key Equations • mass (grams) number of moles

### Objectives Vocabulary Key Equations

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10) What is the molarity of solution made by dissolving 10 g of NaCl into enough water to make 250 mL of saltwater solution? or M L mol L mL g NaCl mol NaCl mL solution g NaCl 0.68 0.68 1 1000 58.44 1 250 10 u u III. Percent Composition and Chemical Formulas 11) A compound analyzed in a chemistry laboratory consists of 5.34 g of carbon, 0.42 g

### Chemistry--Chapter 7: Chemical Quantities

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Chapter Test B A. Matching Match each term in Column B with the correct description in Column A. Write the letter of the correct term on the line. ... CHEMICAL QUANTITIES 10 05\_CTR\_ch10 7/9/04 3:29 PM Page 256. Chapter 10 Review Package - Part 2. Chapter 10 Chemical Quantities. 257 \_\_\_\_\_ 12. What is the molar mass of C. 3. H. 8? a.

### 05 CTR ch10 7/9/04 3:29 PM Page 256 Chapter 10 Review ...

View Notes - Chapter 10 Packet Answer Key from CHEMISTRY Chemistry at Scotch Plains Fanwood Hs. Chapter 10 CHEMICAL QUANTITIES The MOLE Avogadro's Hypothesis Equal volumes of gases (@ same T and p)

### Chapter 10 Packet Answer Key - Chapter 10 CHEMICAL ...

Prentice Hall Chemistry Chapter 10: Chemical Quantities Chapter Exam. Prentice Hall ... Choose your answers to the questions and click 'Next' to see the next set of questions. ... It is shorthand for a substance formula with chemical symbols.

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CHAPTER 10: CHEMICAL QUANTITIES INTRODUCTION In this chapter you will perform many chemical calculations..all of which are based on fundamental principles, such as balanced equations. A balanced equation can provide more information than is apparent at first glance. You can use a balanced equation to help answer such

### CHAPTER 10: CHEMICAL QUANTITIES - ingrum.com

Chemical Quantities THE MOLE AND QUANTIFYING MATTER 10.1 The Mole: A Measurement of Matter Essential Understanding The mole represents a large number of very small particles. Reading Strategy Frayer Model The Frayer Model is a vocabulary development tool. The center of the

### Chemical Quantities - Weebly

The Chemical Quantities Test will be Tuesday, March 10 !! Directions: Use dimensional analysis to perform the following calculations. Show all of your work; including your units and answers in the correct number of significant figures. 1. How many atoms are equal to 4.61 moles of carbon? 03

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